**Additional questions on Solubility product**

**QUESTIONS: For each of the following**

**i)** Write the equation for the equilibrium present in a saturated solution

**ii) Write the expression for Ks**

**iii)** Calculate the solubility (or conductivity) of the following substances in a saturated solution, in mol L–1

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| --- | --- | --- | --- |
| **BaCrO4***K*s(BaCrO4) = 1.2 x 10-10 | **Cu(IO3)2***K*s = 1.3 x 10-12 | **SrF2***K*s = 4.3 x 10-1 | **PbI2***K*s = 8.5 x 10-9 |

**QUESTIONS: For the following questions**

**i)** Write the equation for the equilibrium present in a saturated solution

**ii) Write the expression for Ks for each of the following**

**iii)** Calculate the concentration of the **named ions** in a saturated solution, in mol L–1

|  |  |  |  |
| --- | --- | --- | --- |
| Given Ks of AgCl = 1.8 x 10-10, find the concentration of **Ag+** ions | Given Ks of CuS = 6.0 x 10-37, find the concentration of **S2-** ions | Given Ks of Cu(IO3)2 = 6.9 x 10-8, find the concentration of **Cu2+** ions | Given Ks of Fe(OH)2 = 4.9 x 10-17, find the concentration of **OH**- ions |

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