**ANSWERS: Additional questions on Solubility product**

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| **BaCrO4***K*s(BaCrO4) = 1.2 x 10-10BaCrO4 ⇄ Ba2+ + CrO42-Ks = [Ba2+] [CrO42-]1.2 x 10-10 = x . x1.2 x 10-10 = x2√ 1.2 x 10-10 =x**ANS:** 1.1 x 10-5 mol L-1 | **Cu(IO3)2***K*s = 1.3 x 10-12Cu(IO3)2 ⇄ Cu2+ + 2IO3-Ks = [Cu2+] [IO3-]21.3 x 10-12 = x . 2x21.3 x 10-12 = 4x33√ 1.3 x 10-12 = x **4****ANS:** 6.9 x 10-5 mol L-1 | **SrF2***K*s = 4.3 x 10-10SrF2 ⇄ Sr2+ + 2F-Ks = [Sr2+] [F-]24.3 x 10-10 = x . 2x24.3 x 10-10 = 4x33√ 4.3 x 10-10 = x **4****ANS:** 4.8 x 10-4 mol L-1 | **PbI2***K*s = 8.5 x 10-9*K*s(PbI2) = 8.5 x 10-9PbI2 ⇄ Pb2+ + 2I-Ks = [Pb2+] [I-]28.5 x 10-9 = x . 2x28.5 x 10-9 = 4x33√ 8.5 x 10-9 = x **4****ANS:** 1.2 x 10-3 mol L-1 |

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| **conc of Ag+ ions is****AgCl** ⇄ Ag+ + Cl-Ks = [Ag+] [Cl-]**1.8 x 10 -10 = x . x****1.8 x 10 -19** = x2√ **1.8 x 10 -19 = x****ANS:** 4.2 x 10-10 mol L-1 | **conc of S2- ions is****CuS** ⇄ Cu2+ + S2-Ks = [Cu2+] [S2-]**6.0 x 10 -37 = x . x****6.0 x 10 -37** = x2√ **6.0 x 10 -37 = x****ANS:** 7.7 x 10-19 mol L-1 | **conc of Cu2+ ions is**Cu(IO3)2 ⇄ Cu2+ + 2IO3-Ks = [Cu2+] [IO3-]26.9 x 10-8 = x . 2x26.9 x 10-8 = 4x33√ 6.9 x 10 -8 = x **4****ANS: 2.6 x 10-3** mol L-1 | **conc of OH- ions is**Fe(OH)2 ⇄ Fe2+ + 2OH-Ks = [Fe2+] [OH-]24.9 x 10-17 = x . 2x24.9 x 10-17 = 4x33√ 4.9 x 10-17 = x **4**= 2.3 x 10-6 mol L-1There are 2 x number of OH- ions, so the answer is 4.6 x 10-6 mol L-1 |

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