ANSWERS: Esters and triglycerides

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| **2019** | **Evidence** | **Achievement** | **Merit** | **Excellence** |
| (i)(ii) | Circle around – COO –A hydrolysis reaction uses water to split a large organic moleculeinto smaller organic molecules. Hydrolysis occurs in both acidicand basic conditions (using dilute acid or base). Both acidic and basic hydrolysis require heat under reflux. Both hydrolysis reactions produce the alcohol, glycerol. However, basic hydrolysis will produce the salt of the carboxylic acid, whereas acidic hydrolysis will produce the carboxylic acid.Products from acidic hydrolysis:CH2(OH) – CH(OH) – CH2(OH)3CH3 – (CH2)14 – COOHProducts from basic hydrolysis:CH2(OH) – CH(OH) – CH2(OH)3CH3 – (CH2)14 – COO– | Correctly circles ester group.Describes the hydrolysis reaction.One correct product. | Links acidic OR basic hydrolysis reaction to correct products with structural formulae. | Compares and contrasts acidic andbasic hydrolysis, including ALLcorrect structural formulae ofproducts. |

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| **2018** | **Evidence** | **Achievement** | **Merit** | **Excellence** |
| (i)(ii)(iii)(iv) | ethyl butanoateCH3–CH2–CH2–COO–CH2–CH3Heat under reflux: Diagram 2Increases rate because it is able to be heatedNo loss of products / reactants because they are condensed back into the mixtureIncreases the amount of products / yield because reactants/products are prevented from escapingDistillation could be used to purify the ester (diagram 1).The reaction mixture is heated to the boiling point of the ester which is different from both the alcohol and carboxylic acid reactants. The ester will evaporate from the mixture and enter the condenser where it is cooled back to the liquid to be collected. The ester has thereforebeen separated from the reaction mixture. | • Correct name or structural formula ofethyl butanoate.• Correctly identifies apparatus for BOTHreflux and distillation• Recognises heating under reflux preventsloss of volatile organic compounds / speeds up rate of reaction / increases yield.• Identifies distillation separates ester fromreaction mixture based on its boiling point.ORDescription of distillation process. | Correct diagram for heatingunder reflux TWO advantages and ONE explanation.Correct diagram fordistillation plus process based on different boiling points and separation / purification explained by evaporating, cooling, condensing and collecting. | Fully explains theadvantages of heating under refluxANDFully explains the use of distillation to purify the ester from reaction mixture. |
| **2016** | **Evidence** | **Achievement** | **Merit** | **Excellence** |
| (i)(ii) | One of these groups circled: | * Ester functional group indicated.
* ONE correct product for hydrolysis reaction.
 | * All THREE products correct.
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| **2015** | **Evidence** | **Achievement** | **Achievement with Merit** | **Achievement with Excellence** |
|  (a)(i) | Any one of these groups circled:ass91391Q3a1 | * Functional group correct.
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| (ii) | Bromine water rapidly decolourised from red or orange to colourless in an addition reaction.ORAcidified permanganate rapidly decolourised from purple to colourless in a redox or oxidation or reduction reaction. | * Incomplete description.
 | * Links the observation to the reaction type.
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| (iii) | ass91391Q3a2NaOOC(CH2)7CH=CH(CH2)7CH3 NaOOC(CH2)14CH3  | * ONE product correct.
 | * ALL products correct.
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| **Evidence** | **Achievement** | **Merit** | **Excellence** |
| Ester group: 91391assq1b291391assq1b391391assq1b4Both acidic and basic hydrolysis produce the same alcohol *propan-1,2,3-triol*.In addition, they both require heat / refluxIn contrast, acidic hydrolysis requires H2O / H+ or HCl(*aq*) and produces the carboxylic acid, whereas basic hydrolysis requires H2O / OH– or NaOH(*aq*) and produces the carboxylate ion/salt. | * Correct functional group circled AND named.
* One correct structure.

ORPartial discussion. | * Two unique products linked to the type of hydrolysis.

ORProducts of one reaction correctly linked to the type of hydrolysis and reagent or conditions. | * Comparison and contrast of hydrolysis to give salt / acid and alcohol.

Reagents with conditions and structures correct.  |

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| **2013** | **Evidence** | **Achievement** | **Achievement with Merit** | **Achievement with Excellence** |
|  | * The ester link is hydrolysed in both acid and basic conditions.
* Both produce an alcohol.

Acidic hydrolysis produces an acid and basic hydrolysis produces a base or salt / following hydrolysis in sodium hydroxide, an acid-base reaction occurs to form the sodium salt and water. (No further reaction occurs in acid.) | TWO correct products (name or formula). Identifies hydrolysis. | TWO correct products (name and formula). ANDIdentifies hydrolysis. | ALL products correct (name AND formula).ANDacidic hydrolysis compared and contrasted against basic hydrolysis. |

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