**Atomic structure extra help and questions**

Atoms are the basic building blocks of matter that make up everyday objects. A desk, the air, even you are made up of atoms! Atoms are made up sub-atomic particles which are called electrons, protons and neutrons.  
The centre of an atom is known as the nucleus. The nucleus is where protons and neutrons are found. Electrons are extremely light, they move extremely fast, orbiting the nucleus of an atom in electron shells or clouds. The closest shell/cloud to the nucleus is called the first shell/cloud and holds a maximum of two electrons. For now, you will learn that there are eight electrons orbiting within the other electron shells/clouds.

The numbers of sub-atomic particles for each element are shown on the Periodic table.

**Atomic Number** *(the smallest number)*  
the number of protons   
which is the same as  
the number of electrons

|  |
| --- |
| 2  **He**  4 |

**Mass Number** *(the largest number)*  
the total number of particles in the nucleus; the number of protons   
plus   
the number of neutrons

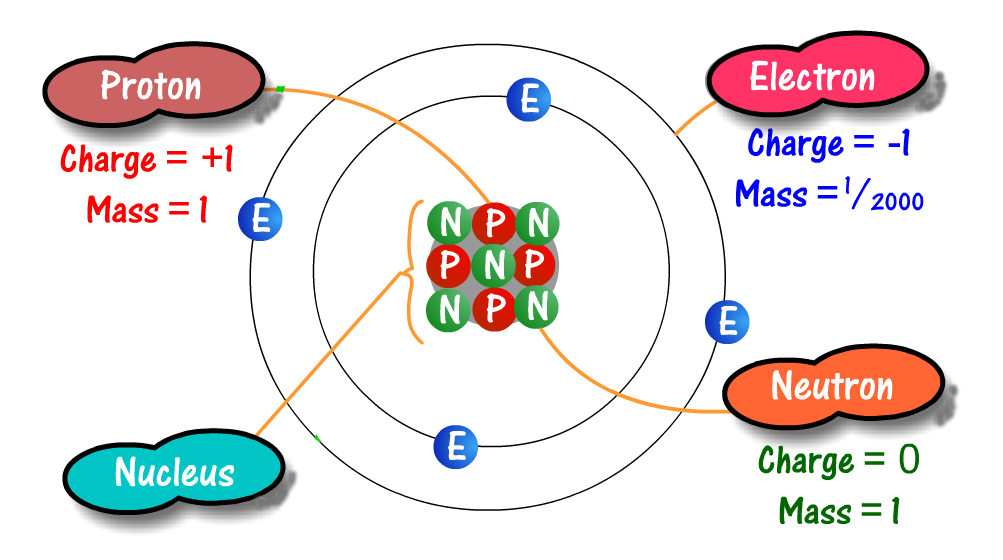
Use the information above to answer the following questions  
**1.** How many electrons, protons and neutrons does an atom of helium have? \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**2.** Draw a neat, labelled sketch of an atom of helium

*please note: You will be provided with a Periodic table in your NCEA Level 1 Science exam. Atomic numbers will be typed on that table. If the mass number is required for a question, then the mass number will be typed within that particular question.*

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**3 (i)** Use at the image above to state the number of protons, electrons and neutrons in this atom

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**(ii)** Use the numbers of protons and neutrons to state the Atomic number and Mass number for this element \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**(iii)** Use both your Periodic table and the image of the atom shown above to write out the name and chemical symbol of this element. \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**4 (i)** The diagram below is of an atom, label the atom with the following words

|  |  |
| --- | --- |
| electron  electron shell  nucleus  proton  neutron |  |

**(ii)** Look at your Periodic table and find out the name and chemical symbol of the element shown above

\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

**5.** Use the periodic table provided in your NCEA Resource booklet to state the number of electrons, and protons in a potassium atom. Also, draw a labelled sketch of the atom

|  |  |  |
| --- | --- | --- |
| \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_  \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ |  | |
| **6. i)** What is the chemical symbol for aluminium?  **ii)** What is the name given to the number next to the letters Al?  **iii)** What does the number tell us?  **iv)** Aluminium is a “neutral” atom, what does neutral mean? | |  |

**7.** The Atomic number for a carbon atom is provided on the periodic table in your resource booklet (it is 6). An exam question may tell you that the Mass number for a carbon atom is 14. Use this information to complete the missing words below using your knowledge of atoms

A carbon atom has an Atomic number of six, so the carbon atom has \_\_\_\_\_ electrons orbiting the nucleus and \_\_\_\_\_\_\_ protons in the nucleus.

The mass number is the total number of particles in the nucleus, so carbon has a total of \_\_\_\_\_ particles in its nucleus. The particles in the nucleus are known as \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. As a carbon atom has 14 particles in the nucleus and 6 of these particles are called protons, then the other \_\_\_\_\_ particles must be neutrons.

Protons have a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charge, neutrons have \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charge and electrons have a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ charge. Electrons are extremely light and move very quickly around the nucleus in \_\_\_\_\_\_\_\_\_\_\_. Only \_\_\_\_\_\_ electrons fit in the first shell/cloud and we have learned (in Year 11) that all of the other shells/clouds fit up to \_\_\_\_\_\_\_ electrons.

**8.** Overleaf is a blank Periodic table (the middle section “Transition metals” is missing) complete each of the following activities on the Periodic table overleaf.

**i)** Label each group of the Periodic table with its correct name or number

**ii)** Label each period of the Periodic table with its correct number

**iii)** Choose 1 group and draw the electrons on the shells for each atom (get the Atomic numbers from your Periodic table)

**iv)** Now write the number of neutrons and protons in the nucleus for the group that you used in ques iii)

**v)** Complete the missing gaps in the following paragraph

Vertical columns of the Periodic table are known as \_\_\_\_\_\_\_\_\_\_\_\_ Horizontal rows of the Periodic table are known as \_\_\_\_\_\_\_\_\_\_\_\_\_ Elements in the same group all have the same number of \_\_\_\_\_\_\_\_\_\_\_\_ in their outer shell *eg I drew the atoms of group \_\_\_\_\_\_\_\_ and noticed that they all have \_\_\_\_ electron(s) in their outer shell*. Elements in the same period all have the same number of \_\_\_\_\_\_\_\_\_\_\_\_ *eg* *I drew the atoms of period \_\_\_\_ and noticed that they all have \_\_\_ shells.*

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| --- | --- | --- | --- | --- | --- | --- | --- |
| 1 | 2 | 13 | 14 | 15 | 16 | 17 | 18 |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |
|  |  |  |  |  |  |  |  |

**ANSWERS**

**6. (i)** Al **(ii)** Atomic number **(iii)** the number of protons which is the same as the number of electrons

**(iv)** “neutral” means that an atom does not have an overall charge because the atom has the same number of positively charged protons (13 protons for aluminium) and negatively charged electrons (13 electrons for aluminium)

**7.** 6, 6,14, protons, neutrons, 6, 6, 8, positive, neutral (or no), negative, electron shells, 2, 8

**8 (v)** groups, periods, electrons, eg... one, one eg…shells, eg…two, two