Calorimetry (Level 3) exam tips: Read these please!

• For water 1mL of water = 1g of water (although this does depend on temperature)

• The heat of reaction = m c ΔT (mass x specific heat capacity of water x change in temperature)

• If needed, the specific heat capacity of water will be given to you as 4.18Jg-1C-1

• We are assuming, of course, that all heat energy lost by the fuel = all heat energy gained by the water

• In reality though, a significant amount of heat energy is lost to the surroundings

Also…”don’t be daft”

Learn the equation as provided above heat of reaction = m c ΔT

Be careful not to use the value for mass of fuel in your equation, m is the mass of the water heated

You MUST be familiar with the practical aspect of calorimetry, if you haven’t carried this out at school there are many online versions

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