Displacement reactions (Level 1) exam tips

• A solid metal combines with a metal ion/compound solution to produce a different metal

ion/compounds and a solid metal.

**OR**

A more reactive metal takes the place of a less reactive metal in a compound.

• Be sure to refer to the Activity Series provided to you in the Resource booklet.

• Link all colours to their species.

• Refer to both the initial and final colours.

• Double check that your equations are balanced.

To achieve with Merit

you must explain displacement reactions in terms of both atoms and ions.

Magnesium is a more reactive metal than copper and so it(magnesium) loses its valence electrons more easily and “forces” the copper ions to accept them

**OR**

Magnesium is a more reactive metal than copperand so will have preference for ending up in its stable ionic/compound form.

Also…”don’t be daft”

Remember that COLOURLESS solutions are just that, never clear!

More reactive elements are higher on the Activity Series (or to the left) in your NCEA Resource booklet

Iron metal is made up of ATOMS with the symbol Fe

Fe(II) or Fe(III) are IONS

Similarly copper METAL is Cu

Cu(II) is the symbol for copper ions

The symbol for potassium is K, not P!!!! (K comes from the medieval Latin word Kalium)

Don’t write an equal’s (=) sign, you MUST have an arrow 🡪

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