Atomic Structure (Level 1) exam tips: Read these please!

• Be able to write out the electron configuration for both atoms and ions

• An atom will either lose or gain electrons to achieve a “stable, full outer electron shell”

• An Ion is “an atom that has lost (*cation*) or gained electron(s) (*anion*)”

• Metal atoms tend to lose electron(s) to form positive ions

• Non-metal atoms tend to gain electron(s) to form negative ions

• **CAT**ion – **CAT**S **p**urr, a **p**ositive ion

• **An**ion – **A** **N**egative ion

• Group number (vertical column) all have the same number of valence (outer) electrons

• Period number (horizontal row) all have the same number of electron shells/energy levels

• RTQ2 be sure to read the question twice

Also…”don’t be daft”

do not use anthropomorphic language such as “happy atoms” or “wishing to..”

© 2018 <https://www.chemical-minds.com>