Describing and Explaining Shapes of molecule (Level 2) exam tips: Read these please!

You must state, describe and explain the following in your answer

• Firstly, locate and name the central atom

• State the number of regions of negative charge (or electron density) around the central atom

• Describe the number of bonding pairs and lone pairs around this central atom

• State that these (*insert number of*) regions of negative charge (or electron density) separate from

each other as much as possible in order to have minimum repulsion

• State the shape and bond angles

Here is a comment from a previous examiners report...*"When explaining the shape of a simple molecule, candidates should demonstrate the understanding that each pair of valence electrons around the central atom repels every other pair of electrons. The molecule is most stable when these repulsive forces are at a minimum, ie. the electron pairs around the central atom are as far apart as possible"*

Also…”don’t be daft”

Don’t write about maximum repulsion, write in terms of MINIMUM repulsion between pairs of electrons

© 2018 <https://www.chemical-minds.com>