Calculating and interpreting Kc

• Kc = products

reactants

• If Kc is large (> 1) then there are more of the product than reactants,

so the right hand/products side is favoured.

• If Kc is small (< 1) then there is more of the reactants than products,

so the left hand/reactants side is favoured.

• The ONLY factor affecting the value of Kc is temperature.  
• If the temperature INcreases the reaction proceeds towards the ENdothermic side (think ENTRANCE

because the endothermic side absorbs the heat energy.

Also…don’t be daft!  
Kc only provides information on how far a reaction proceeds.  
Kc doesn't provide any information on the rate of the reaction.

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