Calculating and explaining equilibrium constant expressions (Level 2)

• K = products (being above the line, known as the numerator)

 reactants (being below the line, known as the denominator)

• If K is large (> 1) then there is more of the product than reactants

 so right hand/products side is favoured

• If K is small (< 1) then there is more of the reactants than products

 so the left hand/reactants side is favoured

• The ONLY factor affecting the value of K is temperature
• If the temperature INcreases the reaction proceeds towards the ENdothermic side (think ENTRANCE

 because the endothermic side absorbs the heat energy

Also…don’t be daft!
K only provides information on how far a reaction proceeds
K doesn't provide any information on the rate of the reaction

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