Identifying types of substances (Level 2) exam tips.

• Be logical here! look out for obvious ones first...

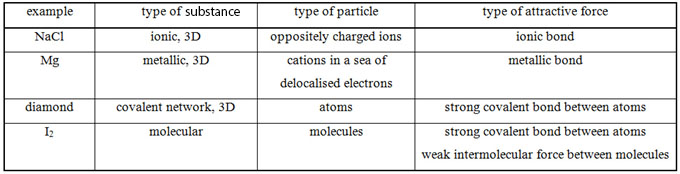
*eg. diamond, graphite and silica or silicon dioxide are covalent network solids.*

*eg metals have metallic bonding.*

*eg metals bonded with non-metals are ionic substances.*

*eg non-metals bonded to themselves are molecular substances.*

• See the table below, a summary of the type of substances, particle and type of bonding/attractive force



Don’t be daft.

Please note that there is no such thing as a molecular bond, no such thing!

Be careful, when describing metallic bonding, it is CATIONS (not atoms!) in a sea of electrons.

Be aware that it is not good enough to state electrostatic attraction as the type of attractive force

because all types of bonding are electrostatic attractions.

© <https://www.chemical-minds.com>