Polymerisation (Level 2) examiners tips

• To form a polymer (*poly means many*)

The double bond between the carbon atoms is broken and new single bonds form, the carbon bonds

from each monomer are able to bond to carbon atoms in other monomers forming a long chin polymer.

• Remember the brackets and “n” indicating that the polymer can be any length or a wiggle line either end

of the polymer

• Be careful when drawing polypropene, a common error is to form a polymer which is in fact polyethene!

A close up of a map

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Also…”don’t be daft”  
The monomer DOES have a DOUBLE bond between the carbon atoms

But **don’t** draw a carbon to carbon double bond in the POLYMER

Remember that every carbon atom must have FOUR bonds from it

Recall the conditions required from your Level 1 Chemistry: high temperature and high pressure with a

catalyst

<https://www.chemical-minds.com>