**Solubility of solids in solutions forming a complex ion**

**2018**

Explain the effect of the following on the solubility of iron(III) hydroxide, Fe(OH)3, in water if potassium

thiocyanate, KSCN, solution is added.

Include relevant equation(s) in your answer. No calculations are necessary.

**2013**

In an experiment, 0.0100 g of Ag2CrO4 in beaker A was made up to a volume of 50.0 mL with

water. In beaker B, 0.0100 g of Ag2CrO4 was made up to a volume of 50.0 mL with 0.100 mol L–1 ammonia solution.



Compare and contrast the solubility of Ag2CrO4 in beaker A and beaker B. *No calculations are necessary.*

**2011**

A saturated solution of zinc hydroxide, Zn(OH)2, contains a small amount of solid Zn(OH)2 at the bottom

of the container. The pH of the solution is increased. Discuss the effect of increasing the pH on the amount

of solid present, and also on the nature and concentration of the species present in the solution.

*No calculations are necessary.*

**2010**

Discuss how the solubility of Ag2CrO4 will change if it is dissolved in the following solution; 0.1 mol L–1 NH3

**No calculations** are necessary.

**2009**

Discuss reasons for the fact that a precipitate of silver chloride dissolves on the addition of excess aqueous ammonia.

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